

Molecular formula	Lewis structure (including any resonance)	Electron Concentrations	Bond angle	Molecular geometry	Name of Molecule or Ion	Polar or nonpolar?
BrCl ₃ 29e ⁻	$ \begin{array}{c} \ddot{\text{Cl}} - \ddot{\text{Br}} - \ddot{\text{Cl}} \\ \\ \ddot{\text{Cl}} \end{array} $	5	90°/120°	trigonal bipyramidal	Bromine trichloride	polar
SO ₃ 24e ⁻	$ \begin{array}{c} \ddot{\text{O}} - \text{S} - \ddot{\text{O}} \\ \quad \text{O} \\ \ddot{\text{O}} \quad \text{O} - \text{S} = \text{O} \end{array} $	3	120°	trigonal planar	Sulfur Trioxide	non polar
SiSe ₂ 16e ⁻	$ \ddot{\text{Se}} = \text{Si} = \ddot{\text{Se}} $	2	180°	linear	silicon diselenide	non polar
ClO ₃ ⁻ 26e ⁻	$ \left[\begin{array}{c} \ddot{\text{O}} - \ddot{\text{Cl}} - \ddot{\text{O}} \\ \\ \ddot{\text{O}} \end{array} \right]^{-} $	4	109.5°	trigonal pyramidal	Chlorate Ion	polar